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CHAPTER 2

ATOMS, MOLECULES, AND IONS

Questions

18. Some elements exist as molecular substances. That is, hydrogen normally exists as H_2 molecules, not single hydrogen atoms. The same is true for N_2 , O_2 , F_2 , Cl_2 , etc.
19. A compound will always contain the same numbers (and types) of atoms. A given amount of hydrogen will react only with a specific amount of oxygen. Any excess oxygen will remain unreacted.
20. The halogens have a high affinity for electrons, and one important way they react is to form anions of the type X^- . The alkali metals tend to give up electrons easily and in most of their compounds exist as M^+ cations. *Note:* These two very reactive groups are only one electron away (in the periodic table) from the least reactive family of elements, the noble gases.
21. Law of conservation of mass: Mass is neither created nor destroyed. The total mass before a chemical reaction always equals the total mass after a chemical reaction.

Law of definite proportion: A given compound always contains exactly the same proportion of elements by mass. For example, water is always 1 g H for every 8 g oxygen.

Law of multiple proportions: When two elements form a series of compounds, the ratios of the mass of the second element that combine with 1 g of the first element always can be reduced to small whole numbers. For CO_2 and CO discussed in Section 2.2, the mass ratios of oxygen that react with 1 g carbon in each compound are in a 2 : 1 ratio.
22. a. The smaller parts are electrons and the nucleus. The nucleus is broken down into protons and neutrons, which can be broken down into quarks. For our purpose, electrons, neutrons, and protons are the key smaller parts of an atom.

b. All atoms of hydrogen have 1 proton in the nucleus. Different isotopes of hydrogen have

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